This Page Is Inserted by IFW Operations and is not a part of the Official Record

BEST AVAILABLE IMAGES

Defective images within this document are accurate representations of the original documents submitted by the applicant.

Defects in the images may include (but are not limited to):

- BLACK BORDERS
- TEXT CUT OFF AT TOP, BOTTOM OR SIDES
- FADED TEXT
- ILLEGIBLE TEXT
- SKEWED/SLANTED IMAGES
- COLORED PHOTOS
- BLACK OR VERY BLACK AND WHITE DARK PHOTOS
- GRAY SCALE DOCUMENTS

IMAGES ARE BEST AVAILABLE COPY.

As rescanning documents will not correct images, please do not report the images to the Image Problem Mailbox.

(12) UK Patent Application (19) GB (11) 2 108 154 A

- (21) Application No 8230680
- (22) Date of filing 27 Oct 1982
- (30) Priority data
- (31) 3142747
- (32) 28 Oct 1981
- (33) Fed Rep of Germany
- (43) Application published 11 May 1983
- (51) INT CL3 C25D 3/56 3/60
- (56) Documents cited C7B 120 701 722 727 737 749 PC U1S 1067 1105 1124 2400 C7B
- (56) Documents cited
 GB 1413180
 GB 1341592
 GB 1246116
 GB 1239862
 GB 0634217
- (58) Field of search C7B C7F
- (71) Applicant
 Maxs AG
 (Switzerland)
 Edisriederstrass 106
 CH-6072 Sachseln OW
 Switzerland
- (72) Inventor
 Albert Greutert
- Albert Greutert
 (74) Agent and/or Address for
 Service
 Fitzpatricks
 Kern House
 61/62 Lincoln's Inn
 Fields
 London WC2B 6EX

(54) Coated heavy metal filters

(57) The invention relates to a filter made of a perforated metal foil coated with a thin layer of an intermetallic compound of tin and nickel.

(12) UK Patent Application (19) GB (11) 2 108 154 A

(21) Application No 8230680

(22) Date of filing 27 Oct 1982

(30) Priority data

(31) 3142747

(32) 28 Oct 1981

(33) Fed Rep of Germany (DE)

(43) Application published 11 May 1983

(51) INT CL3 C25D 3/56 3/60

(56) Documents cited C7B 12O 701 722 727 737 749 PC U1S 1067 1105 1124 2400 C7B

(56) Documents cited

GB 1413180

GB 1341592

GB 1246116

GB 1239862 GB 0634217

(58) Field of search

C7B

C7F (71) Applicant

Maxs AG

(Switzerland)

Edisriederstrass 106

CH-6072 Sachseln OW Switzerland

(72) Inventor

Albert Greutert

(74) Agent and/or Address for

Service

Fitzpatricks

Kern House

61/62 Lincoln's Inn

Fields

London WC2B 6EX

(54) Coated heavy metal filters

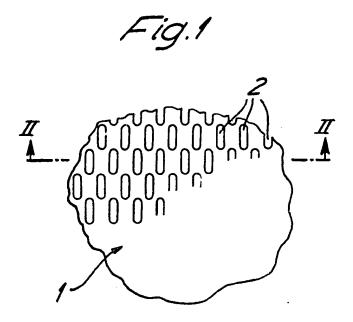
(57) The invention relates to a filter made of a perforated metal foil coated with a thin layer of an intermetallic compound of tin and nickel.

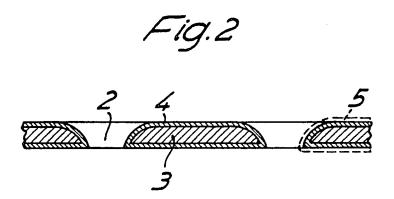
ERRATUM

SPECIFICATION NO 2108155A

Front page, Heading (71) Applicant below Applicant delete whole lines insert Stein Heurtey (France)
Ris Orangis,
France

TE PATENT OFFICE





SPECIFICATION

Perforated metal foil

in them to a small but not negligible extent. This may be a health hazard, more the case of some heavy metals such as nickle, and so it is necessary to inhibit th Accordingly, metal parts of domestic appliances have been given a coating of a gold or platinum or rhodium, which does not enter into solution so readily and/or health hazard. However, it is expensive to use metals of this kind to coat perfora Heavy metals of the same kind suffer from the same disadvantages when the made of them are used as medical filters, for instance, for body fluids. Contact heavy metals and the skin may lead to allergies, for instance, in the case of nick razors. Razor foils of this kind therefore usually have a coating of platinum. 15 It is the object of the invention to provide a perforated metal foil of the kind on the health hazard in contact with ordinary and luxury foods, beverages or the or the skin, yet has a good resistivity and is less expensive than coatings of a promound of tin and nickel. It has been found that a coating of this kind when a coating to the invention, therefore, the foil is coated with a layer of an intercompound of tin and nickel. It has been found that a coating of this kind when a coating to the invention that the concentration of nickel in the coffee is far bel which dissolves, so that the concentration of nickel in the coffee is far bel which dissolves, so that the concentration of nickel in the coffee is far bel which would be a health hazard. To achieve the advantageous properties, the layer may be up to 25 µm thick, between 0.5 and 6µm. The layer may be formed either from a melt or alternatively from an electroplaction. Also, the layer may have in addition a coating of a precious metal, which may for instance in the range of between 0.05 and 0.5 µm, since the intermetallic layer of the nickel foil according to the invention is for coffee filters intermetallic layer of the nickel foil according to the invention is for coffee filters intermetallic layer of the nickel foil according to the inv		•		
gold or piatinum or rhodium, which does not enter into solution so readily and/10 health hazard. However, it is expensive to use metals of this kind to coat perfora Heavy metals of the same kind suffer from the same disadvantages when the made of them are used as medical filters, for instance, for body fluids. Contact heavy metals and the skin may lead to allergies, for instance, in the case of nick razors. Razor foils of this kind therefore usually have a coating of platinum. It is the object of the invention to provide a perforated metal foil of the kind on to a health hazard in contact with ordinary and luxury foods, beverages or the or the skin, yet has a good resistivity and is less expensive than coatings of a procording to the invention, therefore, the foil is coated with a layer of an intercompound of tin and nickel. It has been found that a coating of this kind when a compound of tin and nickel. It has been found that a coating of this kind when a compound of tin and nickel. It has been found that a coating of this kind when a coatine which dissolves, so that the concentration of nickel in the coffee is far bell which would be a health hazard. To achieve the advantageous properties, the layer may be up to 25 μm thick, between 0.5 and βμm. The layer may be formed either from a melt or alternatively from an electrople electrodeposition. Also, the layer may have in addition a coating of a precious metal, which may for instance in the range of between 0.05 and 0.5 μm, since the intermetallic layer of the nickel foil according to the invention is for coffee filters intermetallic layer of the nickel foil according to the invention is for coffee filters intermetallic layer of tin and nickel having a thickness of between 1 and 3μm, a coating being provided of between 0.05 and 0.5 μm thickness. The invention will be described in greater detail hereinafter with reference to drawings of an embodiment, in which Figure 1 is a plan view showing part of a filter foil for a coffee filter, and Figure 2 is a sect		Some metals, when coming into contact with foods, beverages, body fluids or the like dissolve in them to a small but not negligible extent. This may be a health hazard, more particularly in the case of some heavy metals such as nickel, and so it is necessary to inhibit this phenomenon. Accordingly, metal parts of domestic appliances have been given a coating of a metal, such as		
razors. Razor foils of this kind therefore usually have a coating of platinum. It is the object of the invention to provide a perforated metal foil of the kind on to a health hazard in contact with ordinary and luxury foods, beverages or the or the skin, yet has a good resistivity and is less expensive than coatings of a proceeding to the invention, therefore, the foil is coated with a layer of an intercompound of tin and nickel. It has been found that a coating of this kind when incikel which dissolves, so that the concentration of nickel in the coffee is far bell which would be a health hazard. To achieve the advantageous properties, the layer may be up to 25 μm thick, between 0.5 and 6μm. The layer may be formed either from a melt or alternatively from an electroplate electrodeposition. Also, the layer may have in addition a coating of a precious metal, which may for instance in the range of between 0.05 and 0.5 μm, since the intermetallic layer of tin and nickel having a thickness of between 1 and 3μm, a coating being provided of between 0.05 and 0.5 μm thickness. The invention will be described in greater detail hereinafter with reference to a figure 1 is a plan view showing part of a filter foil for a coffee filters, and Figure 2 is a section on the line II-II of Fig. 1. Fig. 1 shows part of a perforated metal foil 1 provided with perforations 2 in having a width of between 50 and 100 μm and a length of approximately 3 mm. The foil comprises a base foil 3 of nickel completely coated by a layer 4 of a comprise of the section of the layer is less than 0.5 μm, to much nickel would distinct metallic compound of tin and nickel. The nickel content of the layer is septimal the tin content 65%. The layer may be applied to the base foil 3 in any required manner, for instar electodeposition from an electroplating bath. The thickness of the layer 4 is between the davantageous effect and should therefore not be considered, since it would unnecessarily long dwell time of the base foil in the bath. An example	10	gold or platinum or rhodium, which does not enter into so health hazard. However, it is expensive to use metals of the Heavy metals of the same kind suffer from the same dismands of them are used as medical filters, for instance, for	olution so readily and/or is not a his kind to coat perforated metal foils. sadvantages when the perforated foils ready fluids. Contact between the	10
 15 It is the object of the invention to provide a perforated metal foil of the kind on ta health hazard in contact with ordinary and luxury foods, beverages or the or the skin, yet has a good resistivity and is less expensive than coatings of a pr According to the invention, therefore, the foil is coated with a layer of an intercompound of tin and nickel. It has been found that a coating of this kind when a common of the content of the kind of the		heavy metals and the skin may lead to allergies, for instar	nce, in the case of nickel folls of dry	
not a health hazard in contact with ordinary and luxury foods, beverages or the or the skin, yet has a good resistivity and is less expensive than coatings of a procording to the invention, therefore, the foil is coated with a layer of an inter compound of tin and nickel. It has been found that a coating of this kind when incikel which dissolves, so that the concentration of nickel in the coffee is far bel which would be a health hazard. To achieve the advantageous properties, the layer may be up to 25 μm thick, between 0.5 and 6μm. The layer may be formed either from a melt or alternatively from an electroplas electrodeposition. Also, the layer may have in addition a coating of a precious metal, which may for instance in the range of between 0.05 and 0.5 μm, since the intermetallic layer of tin and nickel having a thickness of between 1 and 3μm, a coating being provided of between 0.05 and 0.5 μm, since the intermetallic layer of tin and nickel having a thickness of between 1 and 3μm, a coating being provided of between 0.05 and 0.5μm thickness. The invention will be described in greater detail hereinafter with reference to a drawings of an embodiment, in which Figure 2 is a section on the line II-II of Fig. 1. Fig. 1 shows part of a perforated metal foil 1 provided with perforations 2 in having a width of between 50 and 10.0 μm and a length of approximately 3mm. The foil comprises a base foil 3 of nickel completely coated by a layer 4 of a 1 intermetallic compound of tin and nickel. The nickel content of the layer is and the tin content 65%. The layer may be applied to the base foil 3 in any required manner, for instant electodeposition from an electroplating bath. The thickness of the layer is between the advantageous effect and should therefore not be considered, since it would unnecessarily long dwell time of the base foil in the bath. An example of a bath composition for applying the tin/nickel layer by electro follows: The bath is operated with a nickel anode, at a temperature of 70°C and a cur	4-	razors. Razor foils of this kind therefore usually have a con-	metal foil of the kind outlined which is	15
compound of tin and nickel. It has been found that a coating of this kind when instance, to nickel foil coffee filters, provides a reduction of at least 100-fold in nickel which dissolves, so that the concentration of nickel in the coffee is far bel which would be a health hazard. To achieve the advantageous properties, the layer may be up to 25 μm thick, between 0.5 and 6μm. The layer may be formed either from a melt or alternatively from an electropla electrodeposition. Also, the layer may have in addition a coating of a precious metal, which may for instance in the range of between 0.05 and 0.5 μm, since the intermetallic la nickel ensures that only an insignificant or negligible quantity of heavy metal dis 1 have referred use of the nickel foil according to the invention is for coffee filters intermetallic layer of tin and nickel having a thickness of between 1 and 3μm, a coating being provided of between 0.05 and 0.5 μm thickness. The invention will be described in greater detail hereinafter with reference to a drawings of an embodiment, in which Figure 1 is a plan view showing part of a filter foil for a coffee filter, and Figure 2 is a section on the line II-II of Fig. 1. Fig. 1 shows part of a perforated metal foil 1 provided with perforations 2 in having a width of between 50 and 100 μm and a length of approximately 3mm. The foil comprises a base foil 3 of nickel completely coated by a layer 4 of a 40 intermetallic compound of tin and nickel. The nickel content of the layer is apprand the tin content 65%. The layer may be applied to the base foil 3 in any required manner, for instant electodeposition from an electroplating bath. The thickness of the layer is apprand the tin content 65%. The layer may be applied to the base foil 3 in any required manner, for instant electodeposition from an electroplating bath. The thickness of the layer is between the advantageous effect and should therefore not be considered, since it would unnecessarily long dwell time of the base foil in the bath. An examp	15	not a health hazard in contact with ordinary and luxury for the skin, yet has a good resistivity and is less expensive	ods, beverages or the like, body fluids e than coatings of a precious metal.	
 instance, to nickel foil coffee filters, provides a reduction of at least 100-fold in nickel which dissolves, so that the concentration of nickel in the coffee is far bel which would be a health hazard. To achieve the advantageous properties, the layer may be up to 25 μm thick, between 0.5 and 6μm. The layer may be formed either from a melt or alternatively from an electroplate electrodeposition. Also, the layer may have in addition a coating of a precious metal, which may for instance in the range of between 0.05 and 0.5 μm, since the intermetallic la nickel ensures that only an insignificant or negligible quantity of heavy metal disonance in the range of the nickel foil according to the invention is for coffee filters intermetallic layer of tin and nickel having a thickness of between 1 and 3μm, a coating being provided of between 0.05 and 0.5 μm thickness. The invention will be described in greater detail hereinafter with reference to drawings of an embodiment, in which Figure 1 is a plan view showing part of a filter foil for a coffee filter, and Figure 2 is a section on the line II-II of Fig. 1. Fig. 1 shows part of a perforated metal foil 1 provided with perforations 2 in having a width of between 50 and 100 μm and a length of approximately 3mm. The foil comprises a base foil 3 of nickel completely coated by a layer 4 of a 40 intermetallic compound of tin and nickel. The nickel content of the layer is approand the tin content 65%. The layer may be applied to the base foil 3 in any required manner, for instant electodeposition from an electroplating bath. The thickness of the layer 4 is between the layer is limited. A thickness of the layer receeding 6μm is unnece the advantageous effect and should therefore not be considered, since it would unnecessarily long dwell time of the base foil in the bath. An example of a bath composition for applying the tin/nickel layer by electro follows: Tin chloride (SnCl₂.2H₂O) Sivalent tin Nickel chloride (NiCl₂.6H₂O) The bath is opera		According to the invention, therefore, the folials coated	ting of this kind when applied, for	
nickel which dissolves, so that the concentration of nickel in the coffee is far bel which would be a health hazard. To achieve the advantageous properties, the layer may be up to 25 μm thick, between 0.5 and 6μm. The layer may be formed either from a melt or alternatively from an electropla electrodeposition. Also, the layer may have in addition a coating of a precious metal, which may for instance in the range of between 0.05 and 0.5 μm, since the intermetallic la nickel ensures that only an insignificant or negligible quantity of heavy metal dis 1 A preferred use of the nickel foil according to the invention is for coffee filters intermetallic layer of tin and nickel having a thickness of between 1 and 3μm, a coating being provided of between 0.05 and 0.5μm thickness. The invention will be described in greater detail hereinafter with reference to a drawings of an embodiment, in which Figure 1 is a plan view showing part of a filter foil for a coffee filter, and Figure 2 is a section on the line II-II of Fig. 1. Fig. 1 shows part of a perforated metal foil 1 provided with perforations 2 in having a width of between 50 and 100 μm and a length of approximately 3mm. The foil comprises a base foil 3 of nickel completely coated by a layer 4 of a intermetallic compound of tin and nickel. The nickel content of the layer is apparent that the content 65%. The layer may be applied to the base foil 3 in any required manner, for instant electodeposition from an electroplating bath. The thickness of the layer 4 is between 4 the advantageous effect and should therefore not be considered, since it would unnecessarily long dwell time of the base foil in the bath. An example of a bath composition for applying the tin/nickel layer by electro follows: Tin chloride (SnCl ₂ .2H ₂ O) So Sivalent tin Nickel chloride (NiCl ₂ .6H ₂ O) So Nickel Ammonium hydrogen fluoride) Ammonium hydrogen fluoride) Ammonium hydrogen fluoride (NH ₄ F.HF) (Ammonium hydrogen fluoride) Ammonium hydrogen fluoride (The preparation of t	20	instance to nickel foil coffee filters, provides a reduction	of at least 100-fold in the proportion of	20
between 0.5 and 6 μm. The layer may be formed either from a melt or alternatively from an electroplate electrodeposition. Also, the layer may have in addition a coating of a precious metal, which may for instance in the range of between 0.05 and 0.5 μm, since the intermetallic la nickel ensures that only an insignificant or negligible quantity of heavy metal dis A preferred use of the nickel foil according to the invention is for coffee filters intermetallic layer of tin and nickel having a thickness of between 1 and 3μm, a coating being provided of between 0.05 and 0.5μm thickness. The invention will be described in greater detail hereinafter with reference to a drawings of an embodiment, in which Figure 1 is a plan view showing part of a filter foil for a coffee filter, and Figure 2 is a section on the line II-II of Fig. 1. Fig. 1 shows part of a perforated metal foil 1 provided with perforations 2 in having a width of between 50 and 100 μm and a length of approximately 3mm. The foil comprises a base foil 3 of nickel completely coated by a layer 4 of a intermetallic compound of tin and nickel. The nickel content of the layer is approand the tin content 65%. The layer may be applied to the base foil 3 in any required manner, for instant electodeposition from an electroplating bath. The thickness of the layer 4 is between 1 and 1 septiment of the layer is limited. A thickness of the layer exceeding 6μm is unnecessarily long dwell time of the base foil in the bath. An example of a bath composition for applying the tin/nickel layer by electro follows: Tin chloride (SnCl ₂ ·2H ₂ O) Bivalent tin Nickel chloride (NiCl ₂ ·6H ₂ O) Sivalent tin Nickel chloride (NiCl ₂ ·6H ₂ O) Sivalent tin Nickel chloride (NiCl ₂ ·6H ₂ O) The bath is operated with a nickel anode, at a temperature of 70°C and a cur approximately 2.5 A/dm². The rate of deposition is about 1μm/minute. Howew known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee f	20	nickel which dissolves, so that the concentration of nickel in the coffee is far below the level which would be a health hazard.		
The layer may be formed either from a melt or alternatively from an electroplas electrodeposition. Also, the layer may have in addition a coating of a precious metal, which may for instance in the range of between 0.05 and 0.5 μm, since the intermetallic la nickel ensures that only an insignificant or negligible quantity of heavy metal dis 30 A preferred use of the nickel foil according to the invention is for coffee filters intermetallic layer of tin and nickel having a thickness of between 1 and 3μm, a coating being provided of between 0.05 and 0.5μm thickness. The invention will be described in greater detail hereinafter with reference to a drawings of an embodiment, in which Figure 1 is a plan view showing part of a filter foil for a coffee filter, and Figure 2 is a section on the line II-II of Fig. 1. Fig. 1 shows part of a perforated metal foil 1 provided with perforations 2 in having a width of between 50 and 100 μm and a length of approximately 3mm. The foil comprises a base foil 3 of nickel completely coated by a layer 4 of a 1 intermetallic compound of tin and nickel. The nickel content of the layer is approand the tin content 65%. The layer may be applied to the base foil 3 in any required manner, for instant electodeposition from an electroplating bath. The thickness of the layer 4 is between 1 and 10 the layer is less than 0.5 μm, too much nickel would distification of the layer is limited. A thickness of the layer exceeding 6μm is unnece the advantageous effect and should therefore not be considered, since it would unnecessarily long dwell time of the base foil in the bath. An example of a bath composition for applying the tin/nickel layer by electro follows: Tin chloride (SnCl ₂ .2H ₂ O) Bivalent tin Nickel chloride (NiCl ₂ .6H ₂ O) Sivalent tin Nickel chloride (NiCl ₂ .6H ₂ O) Sivalent tin Nickel chloride (NiCl ₂ .6H ₂ O) The bath is operated with a nickel anode, at a temperature of 70°C and a cur approximately 2.5 A/dm². The rate of deposition is about 1μm/minute. Howev known ba			be up to 25 μm thick, preferably	
electrodeposition. Also, the layer may have in addition a coating of a precious metal, which may for instance in the range of between 0.05 and 0.5 μm, since the intermetallic la nickel ensures that only an insignificant or negligible quantity of heavy metal dis 30 A preferred use of the nickel foil according to the invention is for coffee filters intermetallic layer of tin and nickel having a thickness of between 1 and 3μm, a coating being provided of between 0.05 and 0.5μm thickness. The invention will be described in greater detail hereinafter with reference to a drawings of an embodiment, in which Figure 1 is a plan view showing part of a filter foil for a coffee filter, and Figure 2 is a section on the line II-II of Fig. 1. Fig. 1 shows part of a perforated metal foil 1 provided with perforations 2 in having a width of between 50 and 100 μm and a length of approximately 3mm. The foil comprises a base foil 3 of nickel completely coated by a layer 4 of a content of the foil comprises a base foil 3 of nickel completely coated by a layer 4 of a content of the layer is approximately 3mm. The foil comprises a base foil 3 of nickel completely coated by a layer 4 of a content of the layer may be applied to the base foil 3 in any required manner, for instant electodeposition from an electroplating bath. The thickness of the layer 4 is between 1 have 2	25	The layer may be formed either from a melt or alternation	vely from an electroplating bath by	25
Also, the layer may have in addition a coating of a precious metal, which may for instance in the range of between 0.05 and 0.5 μm, since the intermetallic la nickel ensures that only an insignificant or negligible quantity of heavy metal dis 1.5 μm, since the intermetallic layer of tin and nickel having a thickness of between 1 and 3μm, a coating being provided of between 0.05 and 0.5μm thickness. The invention will be described in greater detail hereinafter with reference to a drawings of an embodiment, in which in the figure 1 is a plan view showing part of a filter foil for a coffee filter, and Figure 1 is a plan view showing part of a filter foil for a coffee filter, and Figure 1 is a perforated metal foil 1 provided with perforations 2 in having a width of between 50 and 100 μm and a length of approximately 3mm. The foil comprises a base foil 3 of nickel completely coated by a layer 4 of a 40 intermetallic compound of tin and nickel. The nickel content of the layer is approand the tin content 65%. The layer may be applied to the base foil 3 in any required manner, for instant electodeposition from an electroplating bath. The thickness of the layer 4 is between 1 if the thickness of the layer is less than 0.5 μm, too much nickel would distant advantageous effect and should therefore not be considered, since it would unnecessarily long dwell time of the base foil in the bath. An example of a bath composition for applying the tin/nickel layer by electro follows: Tin chloride (SnCl ₂ .2H ₂ O) Bivalent tin Nickel chloride (NiCl ₂ .6H ₂ O) The bath is operated with a nickel anode, at a temperature of 70°C and a curapproximately 2.5 A/dm². The rate of deposition is about 1μm/minute. Howey known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallic	electrodenosition			
nickel ensures that only an insignificant or negligible quantity of heavy metal dis A preferred use of the nickel foil according to the invention is for coffee filters intermetallic layer of tin and nickel having a thickness of between 1 and 3μm, a coating being provided of between 0.05 and 0.5μm thickness. The invention will be described in greater detail hereinafter with reference to drawings of an embodiment, in which Figure 1 is a plan view showing part of a filter foil for a coffee filter, and Figure 2 is a section on the line II-II of Fig. 1. Fig. 1 shows part of a perforated metal foil 1 provided with perforations 2 in having a width of between 50 and 100 μm and a length of approximately 3mm. The foil comprises a base foil 3 of nickel completely coated by a layer 4 of a intermetallic compound of tin and nickel. The nickel content of the layer is approand the tin content 65%. The layer may be applied to the base foil 3 in any required manner, for instant electodeposition from an electroplating bath. The thickness of the layer 4 is between the thickness of the layer is less than 0.5 μm, too much nickel would distifiction of the layer is limited. A thickness of the layer exceeding 6μm is unnece the advantageous effect and should therefore not be considered, since it would unnecessarily long dwell time of the base foil in the bath. An example of a bath composition for applying the tin/nickel layer by electro follows: Tin chloride (SnCl ₂ ·2H ₂ O) Bivalent tin Nickel chloride (NiCl ₂ ·6H ₂ O) The bath is operated with a nickel anode, at a temperature of 70°C and a cur approximately 2.5 A/dm². The rate of deposition is about 1μm/minute. Howey known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallic.		Also, the layer may have in addition a coating of a precious metal, which may be very thin,		
A preferred use of the nickel foil according to the invention is for coffee filters intermetallic layer of tin and nickel having a thickness of between 1 and 3μm, a coating being provided of between 0.05 and 0.5μm thickness. The invention will be described in greater detail hereinafter with reference to drawings of an embodiment, in which Figure 1 is a plan view showing part of a filter foil for a coffee filter, and Figure 2 is a section on the line II-II of Fig. 1. Fig. 1 shows part of a perforated metal foil 1 provided with perforations 2 in having a width of between 50 and 100 μm and a length of approximately 3mm. The foil comprises a base foil 3 of nickel completely coated by a layer 4 of a intermetallic compound of tin and nickel. The nickel content of the layer is approand the tin content 65%. The layer may be applied to the base foil 3 in any required manner, for instart electodeposition from an electroplating bath. The thickness of the layer 4 is between 1 and 10 μm. If the thickness of the layer is less than 0.5 μm, too much nickel would distingtime of the layer is limited. A thickness of the layer exceeding 6μm is unnecessarily long dwell time of the base foil in the bath. An example of a bath composition for applying the tin/nickel layer by electro follows: Tin chloride (SnCl ₂ .2H ₂ O) Bivalent tin Nickel chloride (NiCl ₂ .6H ₂ O) Solve ammonium bifluoride (NH ₄ F.HF) (Ammonium hydrogen fluoride) Ammonium hydroxide solution (NH ₃ 35% Sp.gr. 880) The bath is operated with a nickel anode, at a temperature of 70°C and a cur approximately 2.5 A/dm². The rate of deposition is about 1μm/minute. Howev known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallic		for instance in the range of between 0.05 and 0.5 μm, si	nter the intermetalic tayer or the and	
intermetallic layer of tin and nickel having a thickness of between 1 and 3μm, a coating being provided of between 0.05 and 0.5μm thickness. The invention will be described in greater detail hereinafter with reference to a drawings of an embodiment, in which Figure 1 is a plan view showing part of a filter foil for a coffee filter, and Figure 2 is a section on the line II-II of Fig. 1. Fig. 1 shows part of a perforated metal foil 1 provided with perforations 2 in having a width of between 50 and 100 μm and a length of approximately 3mm. The foil comprises a base foil 3 of nickel completely coated by a layer 4 of a sintermetallic compound of tin and nickel. The nickel content of the layer is approximated the tin content 65%. The layer may be applied to the base foil 3 in any required manner, for instart electodeposition from an electroplating bath. The thickness of the layer 4 is between the advantageous effect and should therefore not be considered, since it would unnecessarily long dwell time of the base foil in the bath. An example of a bath composition for applying the tin/nickel layer by electro follows: Tin chloride (SnCl ₂ .2H ₂ O) Bivalent tin Nickel chloride (NiCl ₂ .6H ₂ O) So Nickel Ammonium bifluoride (NH ₄ F.HF) (Ammonium hydrogen fluoride) Ammonium hydroxide solution (NH ₃ 35% Sp.gr. 880) The bath is operated with a nickel anode, at a temperature of 70°C and a cur approximately 2.5 A/dm². The rate of deposition is about 1μm/minute. Howev known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallic	30	to A preferred use of the nickel foil according to the invention is for coffee filters, the		
coating being provided of between 0.05 and 0.5μm thickness. The invention will be described in greater detail hereinafter with reference to a drawings of an embodiment, in which Figure 1 is a plan view showing part of a filter foil for a coffee filter, and Figure 2 is a section on the line II-II of Fig. 1. Fig. 1 shows part of a perforated metal foil 1 provided with perforations 2 in having a width of between 50 and 100 μm and a length of approximately 3mm. The foil comprises a base foil 3 of nickel completely coated by a layer 4 of a content of the layer is approximately 3mm. The foil compound of tin and nickel. The nickel content of the layer is approximately 3mm and the tin content 65%. The layer may be applied to the base foil 3 in any required manner, for instart electodeposition from an electroplating bath. The thickness of the layer 4 is between the advantageous effect and should therefore not be considered, since it would unnecessarily long dwell time of the base foil in the bath. An example of a bath composition for applying the tin/nickel layer by electro follows: Tin chloride (SnCl ₂ .2H ₂ O) Bivalent tin Nickel chloride (NiCl ₂ .6H ₂ O) So Nickel Ammonium bifluoride (NH ₄ F.HF) (Ammonium hydrogen fluoride) Ammonium hydroxide solution (NH ₃ 35% Sp.gr. 880) The bath is operated with a nickel anode, at a temperature of 70°C and a cur approximately 2.5 A/dm². The rate of deposition is about 1μm/minute. Howev known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallic	50	intermetallic layer of tin and nickel having a thickness of	layer of tin and nickel having a thickness of between 1 and 3µm, and a gold	
drawings of an embodiment, in which Figure 1 is a plan view showing part of a filter foil for a coffee filter, and Figure 2 is a section on the line II-II of Fig. 1. Fig. 1 shows part of a perforated metal foil 1 provided with perforations 2 in having a width of between 50 and 100 μm and a length of approximately 3mm The foil comprises a base foil 3 of nickel completely coated by a layer 4 of a second that in content 65%. The layer may be applied to the base foil 3 in any required manner, for instart electodeposition from an electroplating bath. The thickness of the layer 4 is between the intervention of the layer is less than 0.5 μm, too much nickel would distributed the advantageous effect and should therefore not be considered, since it would unnecessarily long dwell time of the base foil in the bath. An example of a bath composition for applying the tin/nickel layer by electro follows: Tin chloride (SnCl ₂ .2H ₂ O) Bivalent tin Nickel chloride (NiCl ₂ .6H ₂ O) Solution Tin chloride (SnCl ₂ .2H ₂ O) Solution Tin chloride (SnCl ₂ .2H ₂ O) Solution Tin chloride (SnCl ₂ .2H ₂ O) The bath is operated with a nickel anode, at a temperature of 70°C and a curapproximately 2.5 A/dm². The rate of deposition is about 1μm/minute. Howev known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallic		coating being provided of between 0.05 and 0.5 mm thickness.		
 Figure 1 is a plan view showing part of a filter foil for a coffee filter, and Figure 2 is a section on the line II-II of Fig. 1. Fig. 1 shows part of a perforated metal foil 1 provided with perforations 2 in having a width of between 50 and 100 μm and a length of approximately 3mm. The foil comprises a base foil 3 of nickel completely coated by a layer 4 of a sintermetallic compound of tin and nickel. The nickel content of the layer is approand the tin content 65%. The layer may be applied to the base foil 3 in any required manner, for instart electodeposition from an electroplating bath. The thickness of the layer 4 is beth μm. If the thickness of the layer is less than 0.5 μm, too much nickel would distilifetime of the layer is limited. A thickness of the layer exceeding 6μm is unneces the advantageous effect and should therefore not be considered, since it would unnecessarily long dwell time of the base foil in the bath. An example of a bath composition for applying the tin/nickel layer by electro follows: Tin chloride (SnCl₂.2H₂O) Bivalent tin Nickel chloride (NiCl₂.6H₂O) Tin chloride (NiCl₂.6H₂O) Solution Solution Figure 2 is a section on the layer is limited. The bath is operated with a nickel anode, at a temperature of 70°C and a curapproximately 2.5 A/dm². The rate of deposition is about 1μm/minute. Howey known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallic. 		The invention will be described in greater detail hereinalter with reference to diagrammatic		
Figure 2 is a section on the line II-II of Fig. 1. Fig. 1 shows part of a perforated metal foil 1 provided with perforations 2 in having a width of between 50 and 100 μm and a length of approximately 3mm. The foil comprises a base foil 3 of nickel completely coated by a layer 4 of a 40 intermetallic compound of tin and nickel. The nickel content of the layer is approand the tin content 65%. The layer may be applied to the base foil 3 in any required manner, for instart electodeposition from an electroplating bath. The thickness of the layer 4 is between μm. If the thickness of the layer is less than 0.5 μm, too much nickel would distification of the layer is limited. A thickness of the layer exceeding 6μm is unneces the advantageous effect and should therefore not be considered, since it would unnecessarily long dwell time of the base foil in the bath. An example of a bath composition for applying the tin/nickel layer by electro follows: Tin chloride (SnCl ₂ .2H ₂ O) Bivalent tin Nickel chloride (NiCl ₂ .6H ₂ O) Nickel Ammonium bifluoride (NH ₄ F.HF) (Ammonium hydrogen fluoride) Ammonium hydrogen fluoride) Ammonium hydrogen fluoride The bath is operated with a nickel anode, at a temperature of 70°C and a cur approximately 2.5 A/dm². The rate of deposition is about 1μm/minute. Howey known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallice.	35	Figure 1 is a plan view showing part of a filter foil for a	a coffee filter, and	35
having a width of between 50 and 100 μm and a length of approximately 3mm The foil comprises a base foil 3 of nickel completely coated by a layer 4 of a sintermetallic compound of tin and nickel. The nickel content of the layer is approand the tin content 65%. The layer may be applied to the base foil 3 in any required manner, for instant electodeposition from an electroplating bath. The thickness of the layer 4 is between 1 if the thickness of the layer is less than 0.5 μm, too much nickel would discontent in the savent and should therefore not be considered, since it would unnecessarily long dwell time of the base foil in the bath. An example of a bath composition for applying the tin/nickel layer by electron follows: 50 grams/litre Tin chloride (SnCl ₂ .2H ₂ O) Bivalent tin Nickel chloride (NiCl ₂ .6H ₂ O) 50 Sinckel Ammonium bifluoride (NH ₄ F.HF) (Ammonium hydrogen fluoride) Ammonium hydrogen fluoride) Ammonium hydroxide solution (NH ₃ 35% Sp.gr. 880) 35ml/I The bath is operated with a nickel anode, at a temperature of 70°C and a curapproximately 2.5 A/dm². The rate of deposition is about 1μm/minute. However the seen found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallice	O O	Figure 2 is a section on the line II-II of Fig. 1.		
The foil comprises a base foil 3 of nickel completely coated by a layer 4 of a intermetallic compound of tin and nickel. The nickel content of the layer is approand the tin content 65%. The layer may be applied to the base foil 3 in any required manner, for instart electodeposition from an electroplating bath. The thickness of the layer 4 is between μm. If the thickness of the layer is less than 0.5 μm, too much nickel would distribute of the layer is limited. A thickness of the layer exceeding 6μm is unneces the advantageous effect and should therefore not be considered, since it would unnecessarily long dwell time of the base foil in the bath. An example of a bath composition for applying the tin/nickel layer by electro follows: 50 grams/litre Tin chloride (SnCl ₂ .2H ₂ O) 50 Bivalent tin Nickel chloride (NiCl ₂ .6H ₂ O) 51 Sikel Ammonium bifluoride (NH ₄ F.HF) (Ammonium hydrogen fluoride) Ammonium hydroxide solution (NH ₃ 35% Sp.gr. 880) The bath is operated with a nickel anode, at a temperature of 70°C and a curapproximately 2.5 A/dm². The rate of deposition is about 1μm/minute. Howev known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallice.		Fig. 1 shows part of a perforated metal foil 1 provided	of approximately 3mm	
 40 intermetallic compound of tin and nickel. The nickel content of the layer is apprand the tin content 65%. The layer may be applied to the base foil 3 in any required manner, for instar electodeposition from an electroplating bath. The thickness of the layer 4 is between the layer is limited. A thickness of the layer exceeding 6μm is unneces the advantageous effect and should therefore not be considered, since it would unnecessarily long dwell time of the base foil in the bath. An example of a bath composition for applying the tin/nickel layer by electro follows: 50 grams/litre Tin chloride (SnCl₂.2H₂O) Bivalent tin Nickel chloride (NiCl₂.6H₂O) 50 Nickel Ammonium bifluoride (NH₄F.HF) (Ammonium hydrogen fluoride) Ammonium hydroxide solution (NH₃35% Sp.gr. 880) 35ml/I 60 The bath is operated with a nickel anode, at a temperature of 70°C and a cur approximately 2.5 A/dm². The rate of deposition is about 1μm/minute. Howev known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallice. 		The foil comprises a base foil 3 of nickel completely co	ated by a layer 4 of a monophase	
and the tin content 65%. The layer may be applied to the base foil 3 in any required manner, for instar electodeposition from an electroplating bath. The thickness of the layer 4 is betw μm. If the thickness of the layer is less than 0.5 μm, too much nickel would distall lifetime of the layer is limited. A thickness of the layer exceeding 6μm is unneces the advantageous effect and should therefore not be considered, since it would unnecessarily long dwell time of the base foil in the bath. An example of a bath composition for applying the tin/nickel layer by electro follows: 50 grams/litre Tin chloride (SnCl ₂ .2H ₂ O) Bivalent tin Nickel chloride (NiCl ₂ .6H ₂ O) 50 Sickel Ammonium bifluoride (NH ₄ F.HF) (Ammonium hydrogen fluoride) Ammonium hydrogen fluoride) Ammonium hydroxide solution (NH ₃ 35% Sp.gr. 880) 35ml/l The bath is operated with a nickel anode, at a temperature of 70°C and a cur approximately 2.5 A/dm². The rate of deposition is about 1μm/minute. Howev known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the interrmetallice.	40	intermetallic compound of tin and nickel. The nickel content of the layer is approximately 35%		
electodeposition from an electroplating bath. The thickness of the layer 4 is beth μm. If the thickness of the layer is less than 0.5 μm, too much nickel would distable the advantageous effect and should therefore not be considered, since it would unnecessarily long dwell time of the base foil in the bath. An example of a bath composition for applying the tin/nickel layer by electro follows: 50 grams/litre Tin chloride (SnCl ₂ .2H ₂ O) Bivalent tin Nickel chloride (NiCl ₂ .6H ₂ O) 50 Sivalent tin Nickel chloride (NiCl ₂ .6H ₂ O) 51 Sivalent tin Nickel chloride (NiCl ₂ .6H ₂ O) 52 Sivalent tin Nickel chloride (NiCl ₂ .6H ₂ O) 53 Sp. gr. 880) 35ml/I 50 The bath is operated with a nickel anode, at a temperature of 70°C and a cur approximately 2.5 A/dm². The rate of deposition is about 1μm/minute. Howey known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallice.	. •	and the tin content 65%.		
 μm. If the thickness of the layer is less than 0.5 μm, too much nickel would distable the lifetime of the layer is limited. A thickness of the layer exceeding 6μm is unneces the advantageous effect and should therefore not be considered, since it would unnecessarily long dwell time of the base foil in the bath. An example of a bath composition for applying the tin/nickel layer by electro follows: 50 grams/litre Tin chloride (SnCl₂.2H₂O) Bivalent tin Nickel chloride (NiCl₂.6H₂O) 50 Sivalent tin Ammonium bifluoride (NH₄F.HF) (Ammonium bifluoride (NH₄F.HF) (Ammonium hydrogen fluoride) Ammonium hydroxide solution (NH₃35% Sp.gr. 880) 35ml/I 60 The bath is operated with a nickel anode, at a temperature of 70°C and a cur approximately 2.5 A/dm². The rate of deposition is about 1μm/minute. However known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallice. 	The layer may be applied to the base foil 3 in any required manner, for instance,			
 45 lifetime of the layer is limited. A thickness of the layer exceeding 6μm is unneces the advantageous effect and should therefore not be considered, since it would unnecessarily long dwell time of the base foil in the bath. An example of a bath composition for applying the tin/nickel layer by electro follows: 50 grams/litre Tin chloride (SnCl₂.2H₂O) Bivalent tin Nickel chloride (NiCl₂.6H₂O) Nickel Ammonium bifluoride (NH₄F.HF) (Ammonium hydrogen fluoride) Ammonium hydroxide solution (NH₃35% Sp.gr. 880) 35ml/I 60 The bath is operated with a nickel anode, at a temperature of 70°C and a cur approximately 2.5 A/dm². The rate of deposition is about 1μm/minute. Howev known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallice. 		um. If the thickness of the layer is less than 0.5 um. too	much nickel would dissolve and/or the	
the advantageous effect and should therefore not be considered, since it would unnecessarily long dwell time of the base foil in the bath. An example of a bath composition for applying the tin/nickel layer by electro follows: 50 grams/litre Tin chloride (SnCl ₂ .2H ₂ O) Bivalent tin Nickel chloride (NiCl ₂ .6H ₂ O) 50 Nickel Ammonium bifluoride (NH ₄ F.HF) (Ammonium hydrogen fluoride) Ammonium hydroxide solution (NH ₃ 35% Sp.gr. 880) The bath is operated with a nickel anode, at a temperature of 70°C and a cur approximately 2.5 A/dm ² . The rate of deposition is about 1µm/minute. Howev known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallice.	45	lifetime of the layer is limited. A thickness of the layer exceeding $6\mu m$ is unnecessary to achieve		
An example of a bath composition for applying the tin/nickel layer by electro follows: 50 grams/litre Tin chloride (SnCl ₂ .2H ₂ O) Bivalent tin Nickel chloride (NiCl ₂ .6H ₂ O) 55 Nickel Ammonium bifluoride (NH ₄ F.HF) (Ammonium hydrogen fluoride) Ammonium hydroxide solution (NH ₃ 35% Sp.gr. 880) 35ml/l The bath is operated with a nickel anode, at a temperature of 70°C and a cur approximately 2.5 A/dm ² . The rate of deposition is about 1µm/minute. Howev known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallice.		the advantageous effect and should therefore not be considered, since it would afford an upperessarily long dwell time of the base foil in the bath.		
follows: 50 grams/litre Tin chloride (SnCl ₂ .2H ₂ O) Bivalent tin Nickel chloride (NiCl ₂ .6H ₂ O) 55 Nickel Ammonium bifluoride (NH ₄ F.HF) (Ammonium hydrogen fluoride) Ammonium hydroxide solution (NH ₃ 35% Sp.gr. 880) The bath is operated with a nickel anode, at a temperature of 70°C and a cur approximately 2.5 A/dm². The rate of deposition is about 1µm/minute. Howev known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallice.				
Tin chloride (SnCl ₂ .2H ₂ O) Bivalent tin Nickel chloride (NiCl ₂ .6H ₂ O) 55 Nickel Ammonium bifluoride (NH ₄ F.HF) (Ammonium hydrogen fluoride) Ammonium hydroxide solution (NH ₃ 35% Sp.gr. 880) The bath is operated with a nickel anode, at a temperature of 70°C and a cur approximately 2.5 A/dm². The rate of deposition is about 1μm/minute. Howev known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallice.		follows:	more layer by electing to a	
Tin chloride (SnCl ₂ .2H ₂ O) 50 Bivalent tin 25 Nickel chloride (NiCl ₂ .6H ₂ O) 250 55 Nickel 60 Ammonium bifluoride (NH ₄ F.HF) 40 (Ammonium hydrogen fluoride) Ammonium hydroxide solution (NH ₃ 35% Sp.gr. 880) 35ml/I 60 The bath is operated with a nickel anode, at a temperature of 70°C and a cur approximately 2.5 A/dm ² . The rate of deposition is about 1µm/minute. Howev known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallice.	50			50
Bivalent tin 25 Nickel chloride (NiCl ₂ .6H ₂ O) 55 Nickel Ammonium bifluoride (NH ₄ F.HF) (Ammonium hydrogen fluoride) Ammonium hydroxide solution (NH ₃ 35% Sp.gr. 880) 35ml/I 60 The bath is operated with a nickel anode, at a temperature of 70°C and a cur approximately 2.5 A/dm². The rate of deposition is about 1µm/minute. However, the speed of the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallice.				
Nickel chloride (NiCl ₂ .6H ₂ O) 55 Nickel Ammonium bifluoride (NH ₄ F.HF) (Ammonium hydrogen fluoride) Ammonium hydroxide solution (NH ₃ 35% Sp.gr. 880) 35ml/I 60 The bath is operated with a nickel anode, at a temperature of 70°C and a cur approximately 2.5 A/dm². The rate of deposition is about 1µm/minute. Howeved known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallice.				
 Nickel Ammonium bifluoride (NH₄F.HF) (Ammonium hydrogen fluoride) Ammonium hydroxide solution (NH₃35% Sp.gr. 880) 35ml/I The bath is operated with a nickel anode, at a temperature of 70°C and a cur approximately 2.5 A/dm². The rate of deposition is about 1μm/minute. Howev known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallice. 				
 (Ammonium hydrogen fluoride) Ammonium hydroxide solution (NH₃35% Sp.gr. 880) 35ml/I The bath is operated with a nickel anode, at a temperature of 70°C and a cur approximately 2.5 A/dm². The rate of deposition is about 1μm/minute. Howev known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallice. 	55			55
Ammonium hydroxide solution (NH ₃ 35% Sp.gr. 880) 35ml/I The bath is operated with a nickel anode, at a temperature of 70°C and a cur approximately 2.5 A/dm². The rate of deposition is about 1μm/minute. Howev known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallice.		Ammonium bifluoride (NH ₄ F.HF)	40	
The bath is operated with a nickel anode, at a temperature of 70°C and a cur approximately 2.5 A/dm². The rate of deposition is about 1μm/minute. Howev known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallice.		(Ammonium hydrogen fluoride)	35ml/l	
approximately 2.5 A/dm ² . The rate of deposition is about 1µm/minute. Howev known baths can be used for the preparation of the layer. It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallic			•	60
It has been found that when the metal foil is used as a foil for a coffee filter, nickel dissolving in the case of a nickel base foil, is reduced by the intermetallic	60	approximately 2.5 A/dm². The rate of deposition is about 1μm/minute. However, any other known baths can be used for the preparation of the layer.		
nickel dissolving in the case of a nickel base foil, is reduced by the intermetallic 65 factor of from 100 to 135 as compared with an untreated nickel foil.		It has been found that when the metal foil is used as a	foil for a coffee filter, the proportion of	
00 factor of from 100 to 100 as compared with an unheated nicker for.	0.5	nickel dissolving in the case of a nickel base toil, is reduced to the second second with an untreate	cea by the intermetallic layer 4 by a	65
	65	Tactor of from 100 to 135 as compared with an uniteate	a moror ion.	J U

As shown by the dashed line of Fig. 2, a layer of a precious metal, for instance, a gold layer 5, may additionally be applied to the layer 4. A filter foil of this kind has the advantage that the gold layer may be relatively thin, since it need not act as a barrier layer for nickel but serves merely to ensure that the filter foil has no effect on flavour. Also, the gold layer has to some extent an additional anticorrosive effect. It is also simpler to keep a gold layer clean. 5 The invention is not limited to the use of nickel for the base foil and is applicable to other base metals, which otherwise would form a risk of dissolving to an excessive concentration in foods or the like or where the metal may cause difficulties if it contacts the skin or body fluids. The base foil can be produced in various ways, for instance, by etching or stamping or by 10 deposition from an electroplating bath. 10 **CLAIMS** 1. A perforated metal foil comprising a heavy metal base foil, as a filter for ordinary and luxury foods, beverages or body fluids and as a foil for dry razors, characterised in that the base 15 foil is coated with a layer of an intermetallic compound of tin and nickel. 15 2. A foil according to claim 1, characterised in that the layer has a thickness of up to 25μm. 3. A foil according to claim 1, characterised in that the layer has a thickness of between 0.5 and 6µm. 4. A foil according to claim 1, characterised in that the layer is applied from a melt. 5. A foil according to claim 1, characterised in that the layer is deposited from an 20 electroplating bath. 6. A foil according to claim 1, characterised in that the intermetallic layer is coated with a precious metal. 7. A foil according to claim 6, characterised in that the precious metal coating has a 25 thickness of between 0.05 and 0.5μm. 25 8. A perforated metal foil substantially as hereinbefore described with reference to the

Printed for Her Majesty's Stationery Office by Burgess & Son (Abingdon) Ltd.—1983.
Published at The Patent Office, 25 Southampton Buildings, London, WC2A 1AY, from which copies may be obtained.

accompanying drawings.

7 K . T